In a chemical engineering process, water vapor (H2O) is heated to sufficiently high temperatures that a significant portion of the water dissociates, or splits apart, to form oxygen (O2) and hydrogen (H2):

H2O←→ H2 + 1/2 O2

If it is assumed that this is the only reaction involved, the mole fraction *x* of H2O that dissociates can be represented by

K=x/(1-x) \*√(2p*t*/(2+x))

where *K* is the reaction’s equilibrium constant and *pt* is the total pressure of the mixture. If *pt*= 3 atm and *K* = 0*.*05, determine the value of *x* that satisfies given equation.

* Use graphical model to estimate the value.
* Use Secant method and False Position method to estimate the value for εs=0.5%. Report the number of iterations for each method while achieving the expected result.
* Note: You must write your Secant method and False Position method on separate .m file and you must pass your function as an argument to the method functions. The prototype is given below.
* Secant method (function , 1st initial guess, 2nd initial guess, expected relative approximation error, max iteration)
* False Position method (function , lower bound of the bracket, upper bound of the bracket, expected relative approximation error, max iteration)